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CHAPTER 17, Thermochemistry (continued) 6. In thermochemical calculations, is the direction of heat flow given from the point of view of the system, or of the surroundings? 7. What universal law states that energy can neither be created nor destroyed and can always be accounted for as work, stored potential energy, or heat?

SECTION 17.1 THE FLOW OF ENERGY HEAT AND WORK (pages 505-510)

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I. Name _ Date _ Class _ THERMOCHEMISTRY. SECTION 17.1 THE FLOW OF ENERGY-HEAT AND WORK(pages505-510) This section explains the relationship between energy and heat, and distinguishes between heat capacity and specific heat. ~ Energy Transformations(page505) 1. What area of study in chemistry is concerned with the heat transfers that occur during chemical reactions?thermochemistry.

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You must show all work and setup for this to count as extra credit on your test (+3 points) 15.1 - 15.2 Heat, Calorimetry, and Enthalpy. Make the following conversions: 444 cal to joules . 1.8 kJ to joules . 0.45 kJ to calories . Classify each of these processes as endothermic or exothermic: condensing steam . burning alcohol . evaporating ...

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Chapter 17 Thermochemistry429. Section Review. Objectives. •Explain the relationship between energy, heat, and work. •Distinguish between exothermic and endothermic processes. •Distinguish between heat capacity and specific heat. Vocabulary Key Equations and Relationships. •1 Calorie 1 kilocalorie 1000 calories. •1 J 0.2390 cal and 4.184 J 1 cal.

05 CTR ch17 7/12/04 8:15 AM Page 429 THE FLOW OF ENERGY ...

162 CHAPTER 6: THERMOCHEMISTRY To convert the answer to joules, we write: 101.3 J 0.18 L atm 1L atm = - · x = · w -18 J 6.17 An expansion implies an increase in volume, therefore w must be -325 J (see the defining equation for pressure-volume work.) If the system absorbs heat, q must be +127 J. The change in energy (internal

CHAPTER 6 THERMOCHEMISTRY - Oregon State University

Chapter 17 "Thermochemistry" Tools. Copy this to my account; E-mail to a friend; Find other activities; Start over; Help; Use these activities to learn the vocabulary and major concepts presented in this chapter. A B; calorie: the quantity of heat that raises the temperature of 1 g of pure water by

1°C:

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Chemistry Thermochemistry Assessment Answers Chemistry Chapter 17 Thermochemistry Test - Quizlet The molar mass of ethanol is 4607 g/mol $C_2H_5OH(l) + 3 O_2(g) \rightarrow 2 CO_2(g) + 3 H_2O(g)$ O, 773 -C 30) A 128 g sample of ethanol (C_2H_5OH) is burned in a bomb calorimeter that has a heat capacity of 565

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